

Oxidation/Reduction

Name

Key

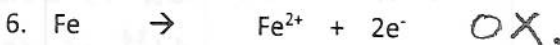
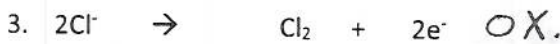
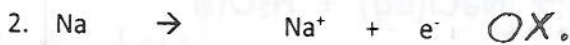
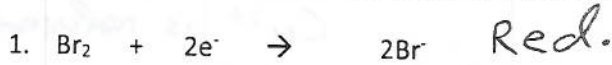
What is oxidation?

Oxidation
Is
Losing electrons
 OIL RIG or LEO the lion says GER

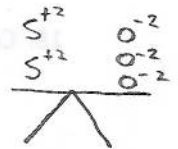
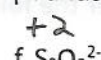
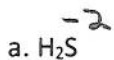
What is reduction?

Reduction
Is
Gaining electrons

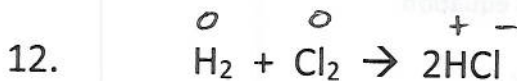
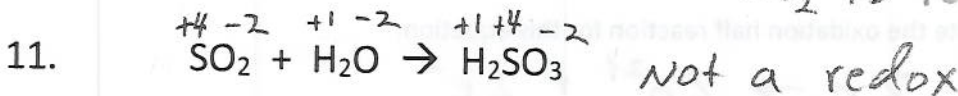
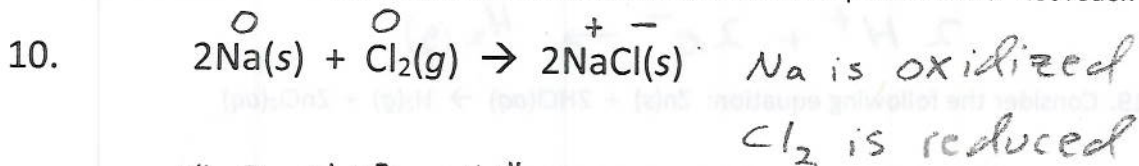
Identify each half-reaction as an oxidation or a reduction half-reaction.



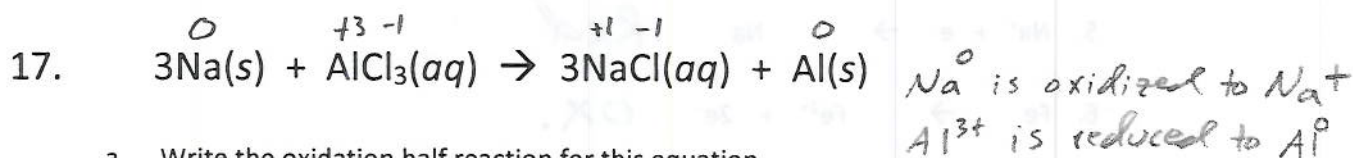
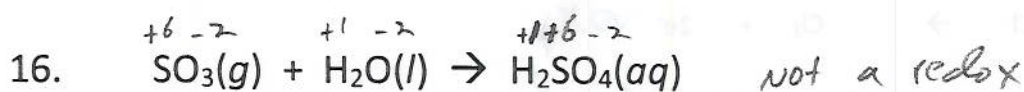
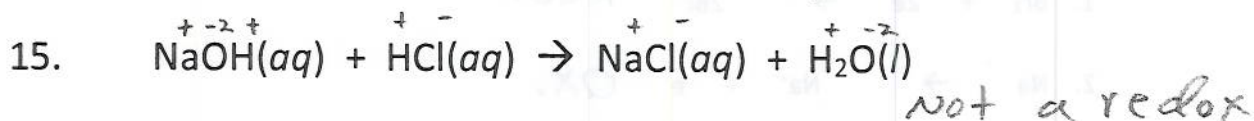
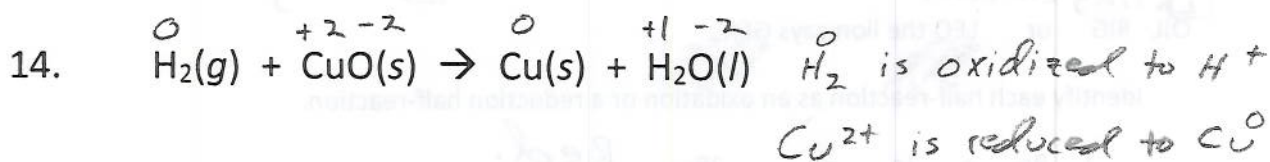
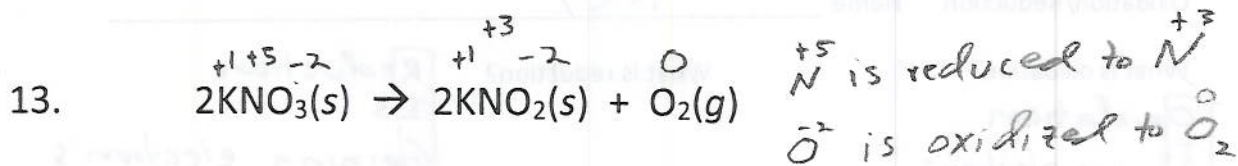
9. Write the oxidation state for sulfur above the "S" in each of the following compounds?



For each equation, determine which element is oxidized and which is reduced. Write the oxidation states above each element. If the equation is not a reduction/oxidation equation write "not redox".



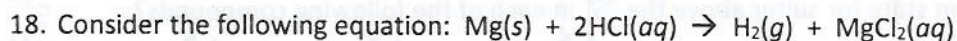
H₂ is oxidized
 Cl₂ is reduced



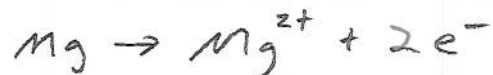
a. Write the oxidation half reaction for this equation.



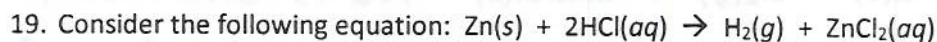
b. Write the reduction half reaction for this equation.



a. Write the oxidation half reaction for this equation.



b. Write the reduction half reaction for this equation.



a. Write the oxidation half reaction for this equation.



b. Write the reduction half reaction for this equation.

